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4296-123

TO THE ASSISTANT COMMISSIONER FOR PATENTS

Transmitted herewith for filing under 35 U.S.C. 111 and 37 C.F.R. 1.53 is the patent application of:

For: **METHOD FOR PRODUCTION OF ACRYLIC ACID AND APPARATUS FOR PRODUCTION OF ACRYLIC ACID**

Enclosed are:

- ☒ Certificate of Mailing with Express Mail Mailing Label No. **EL 512716658 US**
- ☒ 3 sheets of drawings.
- ☐ A certified copy of a application.
- ☒ Declaration ☐ Signed. ☒ Unsigned.
- ☒ Power of Attorney
- ☐ Information Disclosure Statement
- ☐ Preliminary Amendment
- ☒ Other: **Certified copy of Japanese Priority Document No. 11-315914**

CLAIMS AS FILED

For	#Filed	#Allowed	#Extra	Rate	Fee
Total Claims	10	- 20 =	0	x \$18.00	\$0.00
Indep. Claims	2	- 3 =	0	x \$80.00	\$0.00
Multiple Dependent Claims (check if applicable) <input type="checkbox"/>					\$0.00
BASIC FEE					\$710.00
TOTAL FILING FEE					\$710.00

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Dated: **November 3, 2000**



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1. Acknowledgment Post Card;
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3. Patent Application Transmittal Letter (Large Entity)
4. Specification (35 pp.), Claims (2 pp.), Abstract (1 pg.);
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METHOD FOR PRODUCTION OF ACRYLIC ACID
AND APPARATUS FOR PRODUCTION OF ACRYLIC ACID

BACKGROUND OF THE INVENTION

5 Field of the Invention:

This invention relates to a method for the production of acrylic acid, for example, which method comprises preparing the chilled coolant by utilizing the latent heat generated during the gasification of propylene and/or propane as the raw material for the production of acrylic acid and circulating the chilled coolant to the heat exchangers being used in the apparatus for the production of acrylic acid, and to the apparatus for the production of acrylic acid which is suitable for the method mentioned above.

15 Description of Related Art:

Acrylic esters have been finding utility as materials for acrylic fiber copolymers, as emulsions for tackifiers and adhesives, and as raw materials for paints, processed fibers, leather products, and buildings. They have been consequently enjoying growing demands. In the circumstance, the desirability of developing a process which allows use of inexpensive raw materials, permits mass production, and restrains any addition to the environmental pollution has been finding popular recognition. The acrylic acid and the like which are used as raw materials for such acrylic esters are now produced by the reaction of catalytic gas phase oxidation of propylene, for example.

An example of the process for producing acrylic acid by this reaction of catalytic gas phase oxidation will be explained with the aid of Fig. 1 depicting the flow of the raw material for acrylic acid with thick lines and the flow of cooling water with thin lines.

First, the reaction of catalytic gas phase oxidation requires supply of the raw material therefore in the form of the gas as aptly implied by the name. Aliquefied propylene 14 is gasified through the propylene evaporator 3, gaseous propylene from the evaporator is mixed with air 16 for use in the reaction of oxidation, and the resultant mixture is subsequently supplied to the reactor 4. When propylene, for example, is subjected to catalytic gas phase oxidation with the molecular oxygen-containing gas in the presence of the catalyst for the reaction of catalytic gas phase oxidation, the mixed gas containing such by-products as acetic acid in addition to acrylic acid, i.e. the product aimed at is obtained as the reaction product. The reaction product gas formed by the reaction of catalytic gas phase oxidation, therefore, is introduced into the acrylic acid absorbing column 5 and exposed therein to the acrylic acid absorbing solvent, thereby cooled and absorbed, and eventually the liquid containing acrylic acid and such by-products as acetic acid as well is obtained. Then, the solvent in this liquid is separated in the solvent separating column 6 and subsequently distilled and purified in the refining column 7 to separate acrylic acid 18 from by-product containing liquid and obtain the product aimed at.

In this case, a part of the bottom liquid in the acrylic acid absorbing column 5 is introduced into the circulation cooler 9 attached to the acrylic acid absorbing column and cooled by exchanging heat with the liquid coolant supplied from a part of liquid coolant supplying system 1 and then circulated to the acrylic acid absorbing column 5. Incidentally, the acrylic acid absorbing column 5 is generally installed with the absorbing solvent cooler 8 adapted to cool the acrylic acid absorbing solvent 21 for the purpose of

of the reaction system itself tends to lose stability. Further, the concentration of the acrylic acid in the gas supplied to the acrylic acid absorbing column 5 likewise varies and, as a result, polymerization and clogging will happen at the column inside, and the efficiency of absorption will drop in the consequence of the formation of polymer. The decline of the efficiency of absorption lowers the concentration of the acrylic acid in the liquid in which acrylic acid is absorbed emanating from the acrylic acid absorbing column, with the result that the load on the column will rise during the separation of the solvent by the solvent separating column 6.

In the production of the acrylic acid, the acrylic acid absorbing column 5, the solvent separating column 6, the refining column 7, etc. are provided for the purpose of purifying the acrylic acid obtained by the reaction of catalytic gas phase oxidation and these columns are each provided with numerous heat exchangers adapted to enhance the efficiencies of absorption, separation, and purification by utilizing the principle of heat exchange. As concrete examples of the heat exchanger of this nature, the acrylic acid absorbing solvent cooler 8, the condenser 11 attached to the refining column 7, and the condenser 10 attached to the solvent separating column 6 may be installed. The cooling water for use in these heat exchangers is supplied from the liquid coolant supplying system 1 such as a refrigerator and/or cooling tower. Since the temperature of the cooling water varies due to the season, there are cases where the capacity for heat exchange will drop as the rise of cooling water temperature. To prevent this temperature from rising, another electric power is required for cooling the liquid coolant.

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Further, in the case of the solvent separating column 6, for example, the distillate from the top of this column is condensed and cooled by the condenser 10 attached to the solvent separating column. If the efficiency of absorption in the acrylic acid absorbing column 5 falls down, the load exerted to bear on the subsequent solvent separating column 6 will be possibly elevated and the load on the condenser 10 will be consequently elevated. Owing to this rise of load coupled with the fall of the capacity for heat exchange due to the rise of the temperature of the cooling water mentioned above, the distillate from the top of the solvent separating column 6 will not be condensed or cooled sufficiently. Further, the distillate contains acrylic acid, polymerizable by-products, etc. besides the solvent. The solvent separating column 6 is usually operated under vacuum pressure. Under these complex circumstances, insufficient condensation of the distillate will possibly induce such problems as increasing the amount of the polymerizable materials to be scattered in the vacuum equipment connected to the condenser, and cause the vacuum equipment and the piping installed thereto to incur such adverse situation as polymerization and clogging. As a result, it possibly becomes difficult or even impossible to control the operating pressure of the column. Further, these detriments possibly induce the polymerization in the solvent separating column and give rise to such troubles as deterioration of final product quality.

Heretofore, the problem of the polymer formation due to the supply of the steam 17 thereto has been coped with only by periodic removal of the polymer deposited thereon.

SUMMARY OF THE INVENTION

The present inventor has elaborately studied the process

for producing acrylic acid and consequently found that by substituting the liquid coolant for the steam 17 supplied as the heat source for gasifying propylene in the propylene evaporator, it is made possible to allow the reactant gas to be supplied stably and, at the same time, enable the total system for producing acrylic acid to be stabilized drastically, reduce the polymerization and clogging, what is more, allow the electric power heretofore required for the cooler to be decreased by utilizing the chilled coolant obtained during the gasification mentioned above. This invention has been perfected based on this discovery.

To be specific, the tasks imposed on this invention are accomplished by the following items (1) and (2).

(1) A method for the production of acrylic acid or acrolein, characterized by gasifying liquefied propylene and/or propane introduced into an evaporator by supplying a liquid coolant to the evaporator and, at the same time, preparing a chilled coolant by recovering the latent heat of the liquefied propylene and/or propane, subjecting the resultant gasified propane and/or propylene to a catalytic gas phase oxidation reaction thereby preparing a gas containing acrylic acid or acrolein, and using said chilled coolant in heat exchangers attached to the apparatus for the production acrylic acid or acrolein.

(2) An apparatus for the production of acrylic acid or acrolein, comprising means for gasifying liquefied propylene and/or propane introduced into an evaporator by supplying a liquid coolant to said evaporator and, at the same time, preparing a chilled coolant by recovering the latent heat of the liquefied propylene and/or propane, means for subjecting the resultant gasified propylene and/or propane to a catalytic gas phase oxidation reaction thereby preparing

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a gas containing acrylic acid or acrolein, and means for using said chilled coolant in heat exchangers attached to the apparatus for the production of acrylic acid or acrolein.

According to this invention, since the latent heat is recovered by liquid coolant, the gasification of propylene and/or propane can be effected stably and the supply of the reactor with propylene can be attained stably as compared with the conventional method utilizing steam. As a result, the absorption and the distillation in the acrylic acid absorbing column, solvent separating column, refining column, etc. can be stabilized and the final product can be improved in quality. At the same time, by using the chilled water prepared by the method mentioned above in the heat exchangers in the process of acrylic acid production, it is made possible to decrease the energy consumption for cooling.

BRIEF DESCRIPTION OF THE DRAWINGS

Fig. 1 is a diagram illustrating an example of the conventional process for production of acrylic acid by the reaction of catalytic gas phase oxidation, with the flows of raw materials for acrylic acid indicated in thick lines and those of cooling water in thin lines.

Fig. 2 is a diagram illustrating a preferred working example of this invention for producing acrylic acid by a two-step reaction using water as the liquid coolant and propylene as the raw material.

Fig. 3 is a diagram illustrating a preferred working example of this invention for producing acrylic acid by using brine as the liquid coolant.

DESCRIPTION OF THE PREFERRED EMBODIMENTS

The first aspect of this invention consists in a

method for the production of acrylic acid or acrolein, characterized by gasifying liquefied propylene and/or propane introduced into an evaporator by supplying a liquid coolant to the evaporator and, at the same time, preparing a chilled
5 coolant by recovering the latent heat of the liquefied propylene and/or propane, subjecting the resultant gasified propane and/or propylene to a catalytic gas phase oxidation reaction thereby preparing a gas containing acrylic acid or acrolein, and using said chilled coolant in heat exchangers
10 attached to the apparatus for the production acrylic acid or acrolein.

It has been heretofore customary to use steam 17 in gasifying liquefied propylene and/or propane. Since the steam has high energy, even a slight fluctuation of its pressure
15 or supplied amount varies the composition of the reactant gas supplied to the reactor for catalytic gas phase oxidation and deprives the catalytic gas phase oxidation reaction itself of its stability. This invention, however, is enabled by using the liquid coolant in the place of the steam 17 to stabilize
20 the gasification itself and also stabilize the whole reaction system and the whole of the subsequent steps for producing acrylic acid as well. The drain resulting from the use of steam 17 has been heretofore reused barely as the boiler feed water for the purpose of utilizing sensitive heat. Owing to
25 the use of the liquid coolant, however, this invention enables the chilled coolant to be obtained in a large amount in consequence of the gasification and allows the reduction of the electric power consumed in the heat exchangers in the acrylic acid process by using the obtained chilled coolant.
30 The term "liquid coolant" as used in this invention refers to the cooling agent which is used in gasifying the liquefied propylene and/or propane destined to serve as the raw material

For a start, the liquid coolant supplied from the liquid coolant supplying system 1 is introduced to the propylene evaporator 3 via line L1 to gasify the liquefied propylene. The evaporator 3 is connected with the piping for introducing the liquid propylene and has empty space for introducing the liquid coolant, and enables gasification of liquefied propylene by the principle of heat exchanger, with the result that the liquid coolant will be converted into the chilled coolant.

As the liquid coolant, water or brine may be used. As the water, industrial water, city water, or other water may be used. The brine is known in numerous types such as ethylene glycol, methanol, and other aqueous solutions. This invention allows use of the brine as the known liquid coolant to be circulated to the heat exchangers. The brine is at an advantage in retaining the liquid state at lower temperature than water and allowing the lowest temperature for the liquid state to be adjusted in accordance with the composition thereof. The choice of water or brine as the liquid coolant and the question as to the kind of composition for the brine to be used may be properly decided to suit the kind of use and the purpose of use.

The method for production according to this invention has no particular restriction for the temperature of the liquid coolant to be supplied to the evaporator 3. This is because the liquid coolant is capable of stabilizing the reaction system and reducing the power for cooling so long as it is in a liquid state at the time that it is introduced into the evaporator 3. In the gasification of the liquefied propylene and/or propane, the temperature is in the range of 0 - 50°C, preferably 5 - 40°C, and particularly preferably 5 - 35°C. The reason for this range of temperature will be described

in detail below.

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First, the vaporizing temperature of the liquefied propylene 14 in the evaporator 3 varies with the saturated vapor pressure of propylene in the evaporator. Since this
5 saturated vapor pressure can be controlled by the operating pressure of the liquefied propylene 14 introduced into the evaporator 3, the vaporizing temperature may be set at an arbitrary level by adjusting the operating pressure of the liquefied propylene 14. The gasification of the liquefied
10 propylene 14 requires the vaporizing temperature to be lowered below the temperature of the liquid coolant to be supplied, and this will inevitably entail the drop internal pressure, that is, the operating pressure of the evaporator. The drop of the operating pressure, however, results in increasing
15 the specific volume of the gasified propylene and consequently requires the volume of the evaporator 3 itself to be enlarged for the purpose of retaining the rate of vaporization at a proper level in the evaporator 3. Specifically, when the temperature of the liquid coolant falls below 0°C, it becomes
20 necessary to enlarge the volume of the evaporator 3 and, as a natural consequence, to enlarge the pipe lines through which the gasified propylene flows. During the stop of the operation or during an erroneous operation, the interior of the evaporator 3 is exposed to high pressure because the
25 temperature of the propylene in the evaporator 3 rises to the same temperature as the liquid coolant to be supplied. The evaporator 3, therefore, must be designed with excessive pressure resist ability.

30 Conversely, when the temperature of the liquid coolant is high, the volume of the evaporator 3 may be reduced due to the rise of operating pressure in compliance with the above mentioned theory and practice. When the liquefied propylene

4 and enabled consequently to put to circulation part of the bottom of the column and, at the same time, supply the cooled acrylic acid absorbing solvent 21 from the top of the column, with the result that the temperature of the top of the column and the efficiency of acrylic acid absorption will be maintained at respectively prescribed levels. As the liquid coolant used in the circulation cooler 9, this invention allows use of the chilled coolant prepared as described above. Incidentally, the temperature of the chilled coolant prepared in the evaporator and/or adjusted by the liquid coolant thermocontroller is not always required to be equal to the temperature of the chilled coolant which is supplied to the coolers 8 and 9, etc. The chilled coolant to be circulated to the circulation cooler 9, for example, does not need to be limited to the chilled coolant which is cooled by the refrigerator 13. The chilled coolant withdrawn via the branch from the line 2 extending from the evaporator 3 through the refrigerator 13 may be used instead. When the chilled coolant is circulated to the circulation cooler 9, the temperature thereof is properly in the range of 0 - 40°C, preferably 5 - 35°C. The chilled coolant which has undergone heat exchange in the circulation cooler 9 is introduced via the line 5 and the chilled coolant which has undergone heat exchange in the cooler 8 is introduced via the line 4 and they are mixed. It is preferable to mix the chilled coolant which has different temperature and to return to resultant mixed chilled coolant to the liquid coolant supply system 1 via the line 6 and reuse as liquid coolant.

The chilled coolant emanating from the refrigerator 13 can be advanced through the branch from the line 3 and circulated to the condenser 10 attached to the solvent separating column 6 and used therein as the liquid coolant

for heat exchange. As the chilled coolant to be used for this purpose, that which has been cooled by the refrigerator 13 proves favorable. When the chilled coolant is circulated to the condenser 10, the temperature thereof is properly in the range of 0 - 35°C, preferably 5 - 30°C. Incidentally, in order to prepare the chilled coolant which has appropriate temperature as mentioned above, the condenser 10 may be provided with the thermocontroller. The chilled coolant which has undergone heat exchange may be advanced from the condenser 10 through the line 7 and joined to the flow in the line 6, again returned to the liquid coolant supplying system 1 and reused as the liquid coolant.

The process for producing acrylic acid usually involves the acrylic acid refining column 7 in addition to the acrylic acid absorbing column and the solvent separating column mentioned above. When the acrylic acid refining column 7 is involved in the process, the chilled coolant mentioned above may be used as the liquid coolant for the heat exchanger attached to the acrylic acid refining column 7 such as, for example, the condenser 11 attached to the refining column. When the chilled coolant is circulated to such a heat exchanger, the temperature thereof is properly in the range of 20 - 35°C, preferably 20 - 30°C. The vapor of acrylic acid is distilled from the acrylic acid refining column 7 through the top thereof and subsequently cooled by the condenser 11 attached to the refining column to obtain acrylic acid. Meanwhile, the liquid 19 containing the by-product is recovered through the bottom of the column. Incidentally, the liquid coolant which has utilized in the condenser 10 attached to the solvent separation column or the condenser 11 attached to the refining column is passed through the line 7 and the line 8 and mixed with the flow through the line 6 and returned to the liquid coolant

supplying system 1 and reused as the liquid coolant. When the apparatus for the production of acrylic acid mentioned above is connected to the methacrylic acid and/or (meth)acrylic esters plants 12, the liquid coolant mentioned above may be supplied as the liquid coolant for heat exchange in the plant 12 and the liquid coolant which has undergone this heat exchange advanced through the line 9, combined with the flow through the line and returned again to the liquid coolant supplying system 1 and reused as the liquid coolant.

One example of the chilled coolant used in the process for producing acrylic acid by the series of operations resorting to the reactor 4, acrylic acid absorbing column 5, solvent separating column 6, and refining column 7 has been installed. This invention allows the chilled coolant to be used in such heat exchangers installed in the plants other than the plant for producing acrylic acid and the plants further connected to the acrylic acid plant such as, for example, the methacrylic acid and/or (meth)acrylic esters plants 12. When the amount of the heat of the chilled coolant calculated from the amount of the chilled coolant obtained by gasification is in excess of the total amount of the heat required for cooling in the heat exchangers attached to the process for producing acrylic acid, when the amount of the heat of chilled coolant is in excess amount because the chilled coolant has been further cooled by the refrigerator 13, and when the chilled coolant usable for cooling is in excess amount because it has been used only in a part of the heat exchangers involved in the process for producing acrylic acid, these excess of the chilled coolant can be effectively utilized in the plants mentioned above instead of being wasted. Moreover, the use of the chilled coolant results in not only reduction of energy consumption for cooling but also stabilizing the process of

being used therein, forwarded via the line 10 and combined with the flow in the line 6 and circulated to the liquid coolant supplying system 1.

Now, the supply route for the propylene and/or propane emanating from the evaporator 3 will be described. The propylene, for example, which has been gasified by the evaporator 3 is mixed with the inert gas and/or air through the inert gas supply route and the air supply route which are omitted from illustration herein and the reactant gas consequently formed is supplied to the reactor 4. Further, this invention can adjust the amount of the propylene and/or propane 14 to be gasified by properly controlling the amount of the liquid coolant introduced into the evaporator 3 as described above. That is, this invention can stabilize conversion and the yield in the step for producing acrylic acid by adjusting the amount of the liquid coolant to be introduced. The adjustment of the amount of the gasified propylene and/or propane supplied from the evaporator 3 and introduced into the reactor 4 may be attained by adjusting the gas flow rate using gas flow rate controller 25 installed in the line.

In the present invention, the reaction of catalytic gas phase oxidation of propylene with molecular oxygen containing gas can be carried out under the conditions adopted by the heretofore known method. In the case of propylene, for example, the reactant gas used for the reaction has appropriate propylene concentration in the range of 3 - 15 vol. %, the ratio of molecular oxygen containing gas to propylene in the range of 1 - 3, and comprises nitrogen, steam, carbon oxides, and propane, etc.

As the source for the molecular oxygen containing gas, air is advantageously used. Oxygen-enriched air and purified

oxygen, optionally, may be used instead. This source is supplied by the one-pass mode or the recycling mode. Commendably, the reaction temperature is in the range of 250°C - 450°C, the reaction pressure in the range of absolute pressure
5 0.1 - 0.5 MPa, and the space velocity in the range of 500 - 3000 h⁻¹ (STP).

Though the reactor 4 does not need to be particularly discriminated on account of the kind, the shell-and-tube type reactor may be preferably used because it is effective for
10 heat exchange.

As the catalyst with which the reactor 4 is to be packed for the production of acrylic acid by two-stage catalytic gas phase oxidation reaction, the oxidation catalyst which is generally used for the production of acrolein by the gas
15 phase oxidation reaction of reactant gas can be used as the former-stage catalyst. By the same token, the latter-stage catalyst does not need to be particularly discriminated. The oxidizing catalyst which is generally used in the production of acrylic acid by the gas phase oxidation of the reaction
20 gas mainly containing the acrolein obtained by the former stage in the method of the two-stage catalytic gas phase oxidation can be used, for example.

As concrete examples of the former-stage catalyst, the compositions represented by the general formula,
25. Mo_a-Bi_bFe_c-A_d-B_e-C_f-D_g-O_x (wherein Mo, Bi, and Fe respectively denote molybdenum, bismuth, and iron, A denotes at least one element selected from the group consisting of nickel and cobalt, B denotes at least one element selected from the group consisting of alkali metals and thallium, C denotes at least
30 one element selected from the group consisting of phosphorus, niobium, manganese, cerium, tellurium, tungsten, antimony, and lead, D denotes at least one element selected from the

group consisting of silicon, aluminum, zirconium, and titanium, and O denotes oxygen, and a, b, c, d, e, f, g, and x respectively denote the atomic ratios of Mo, Bi, Fe, A, B, C, D and O which fall respectively in $b = 0.1 - 10$, $c = 0.1 - 10$, $d = 2 - 20$, $e = 0.001 - 5$, $f = 0 - 5$, and $g = 0 - 30$ when $a = 12$ is fixed, and x denotes the value determined by the states of oxidation of the relevant elements) may be cited.

As concrete examples of the latter-stage catalyst, the compositions represented by the general formula, $Mo_a-V_b-W_c-Cu_d-A_e-B_f-C_g-O_x$ (wherein Mo denotes molybdenum, V denotes vanadium, W denotes tungsten, Cu denotes copper, A denotes at least one element selected from the group consisting of antimony, bismuth, tin, niobium, cobalt, iron, nickel, and chromium, B denotes at least one element selected from the group consisting of alkali metals, alkaline earth metals, and thallium, C denotes at least one element selected from the group consisting of silicon, aluminum, zirconium, and cerium, and O denotes oxygen, a, b, c, d, e, f, g, and x respectively denote the atomic ratios of Mo, V, W, Cu, A, B, C, and O which fall in the ranges of $b = 2 - 14$, $c = 0 - 12$, $d = 0.1 - 5$, $e = 0 - 5$, $f = 0 - 5$, and $g = 0 - 20$ when $a = 12$ is fixed, and x denotes the value determined by the states of oxidation of the relevant elements) may be cited.

The catalyst does not need particularly to discriminate the shape thereof but may be in the shape of spheres, circular columns, cylinders, etc. It may be formed by such methods as carrier molding, extrusion molding, and tablet molding. The catalyst of the type having a catalytic substance deposited on a refractory carrier is particularly useful.

The production of the acrylic acid is accomplished by introducing the reactant gas under the conditions mentioned

above into the former-stage catalyst layer and then introducing the formed gas mainly containing acrolein either directly or after having added air and steam into the latter-stage catalyst layer.

Specifically, for the purpose of forming the acrylic acid, the reaction tubes in the reactor 4 are packed with the oxide catalyst (latter-stage catalyst) and the acrolein-containing gas obtained by the former-stage reaction or the mixed gas prepared, when necessary, by adding air, oxygen, or steam to the gas is supplied to the reaction tubes at the reaction temperature (the temperature of the catalyst in the reactor) in the range of 100 - 380°C, preferably 150 - 350°C, and the space velocity in the range of 300 - 5,000 hr⁻¹ (STP) and subjected to the latter-stage reaction to produce the acrylic acid.

Incidentally, the formation of the acrolein-containing gas does not need to be limited to such two-stage reaction of catalytic gas phase oxidation as described above but may be attained by using different reactors respectively for the two stages involved and performing the reactions in two or more stages. In this case, the acrolein-containing gas is formed by one-stage reaction of catalytic gas phase oxidation. Similarly, propane may be used in the place of propylene and the latent heat may be recovered and utilized as in an apparatus for producing acrylic acid.

The reactor 4 generates a great deal of heat of reaction. As the heat medium for circulation in the reactor for the purpose of removing this heat of reaction, any of the heat medium heretofore known to the art may be used. As concrete examples of the heat medium for circulation in the reactor, molten salts, niter, and phenyl ether type heat medium which is the Dowtherm type organic heat medium may be cited. The

reactor heat medium which has the temperature thereof elevated by removing the heat of the reaction tubes is cooled in the heat exchanger attached to the reactor.

Then, the acrylic acid-containing gas 20 emanating from the reactor 4 is supplied to the acrylic acid absorbing column 5. In the absorbing column 5, the absorbing solvent 21 is supplied thereto via the absorbing solvent cooler 8 and utilized therein to prepare the acrylic acid-containing solution. The waste gas 20 emanating from the top of the absorbing column 5 is either discarded or recycled. By using the cooled acrylic acid absorbing solvent 21, it makes possible to improve the efficiency of acrylic acid absorption in the acrylic acid absorbing column 5.

The absorbing solvent 21 in the absorbing column 5 may be used any of the known absorbing solvents such as, for example, water, organic acid-containing waters, and the solvents which have high boiling point such as inactive hydrophobic organic solvents (represented by diphenyl ether and diphenyls). When water is used as absorbing solvent 21, for example, the aqueous acrylic acid solution to be obtained from the acrylic acid absorbing column 5 under the generally employed conditions for the production of acrylic acid generally contains acrylic acid in the range of 50 - 80 mass %, acetic acid in the range of 1 - 5 mass %, and water in the range of 20 - 40 mass %. These concentration are varied by the operating conditions of the oxidation reaction in the reactor or the absorbing column. Incidentally, the temperature at the top of the acrylic acid absorbing column is preferred to be in the range of 40 - 70°C.

Subsequently, the acrylic acid-containing solution is supplied to the solvent separating column 6 and separated therein into acrylic acid and the solvent. The distillate

emanating from this column from the top thereof is condensed and cooled by the condenser 10 attached to the column. In case the solvent which has lower boiling point than acrylic acid such as water is used as the absorbing solvent, the acrylic acid is withdrawn through the bottom of the column. In case the solvent has higher boiling point mentioned above, the acrylic acid is withdrawn through the middle of the column. In any cases, the column can be operated under the heretofore known conditions to be employed when such solvents are used.

Even when the solvent separating column is required to be followed by an acrylic acid refining column in the sequence of process for the purpose of obtaining acrylic acid as a product, the operation of the refining column may be performed under the conditions heretofore known to the art.

When the refining column is required, this column is additionally provided with the condenser 11 which fulfills the function of condensing and cooling the distillate emanating from the top of the column. As the liquid coolant used in this condenser 11, the chilled coolant which is prepared by the gasification mentioned above may be used. The liquid coolant which has been used in the condenser 11 as described above may be circulated to the liquid coolant supplying system 1. Incidentally, Figs. 2 and 3 depict the processes which use the solvent having a lower boiling point than acrylic acid as the acrylic acid absorbing solvent and recover the acrylic acid-containing solution through the bottom of the solvent separating column 6.

The method for the production of acrylic acid which includes the steps for producing acrylic acid by absorbing acrylic acid by using the solvent having high boiling point and subjecting the acrylic acid obtained from the solvent separating column 6 to further refinement and a step for

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is no longer required. Particularly, however, the brine is capable of retaining the lower temperature than water. In case the apparatus which produces (meth)acrylic acid and/or (meth)acrylic ester and involves cyclic use of the chilled coolant, for example, has part thereof suspended from operation, there are cases where the temperature of the liquid coolant falls to extremely low level. In this case, therefore, it is preferable to make the liquid coolant temperature controlled in the range of 0-50°C by using the liquid coolant heater 2 and then introduce it into the evaporator 3 through the line 1-1. When the brine is used, for example, by supplying brine having an ethylene glycol concentration of 30 mass % at the temperature in the range of 0 - 30°C to the evaporator 3, it makes possible to obtain the chilled coolant having the temperature in the range of -5 - 25°C.

When the plant is provided with lines so installed that the chilled coolant used in any of the heat exchangers incorporated in the plant may be recovered in the liquid coolant tank 27, it makes possible to store the used chilled coolant in the liquid coolant tank 27 and transfer the liquid coolant from the tank using the liquid coolant transfer pump 28 which is installed in the line 11 connected to the tank. When the liquid coolant tank 27 is disposed as described above, the pressure accumulating in the line for the liquid coolant transfer and the non-condensable gas possibly leaking in a minute amount into the liquid coolant may be separated. Particularly since the brine can be a chilled coolant of a lower temperature than water, the liquid coolant may be circulated only partly to the evaporator 3 and used as the chilled coolant instead of being wholly circulated to the evaporator 3. For example, part of the liquid coolant flowing from the liquid coolant transfer pump 28 through the line

gasified propylene and/or propane to a catalytic gas phase oxidation reaction thereby preparing a gas containing acrylic acid or acrolein, and means for using said chilled coolant in heat exchangers attached to the apparatus for the production of acrylic acid or acrolein.

The acrylic acid is produced through acrolein by the reaction of catalytic gas phase oxidation of propylene and/or propane. The apparatus for production according this invention recovers the latent heat of liquefied propylene and/or propane and stabilize the process for producing acrylic acid or acrolein as well and, consequently, produces acrylic acid of high quality and decreases the amount of electric energy consumed for cooling in the equipments, for examples, heat exchangers of the process of the production.

In the apparatus of this invention, as the "means for gasifying liquefied propylene and/or propane introduced into the evaporator by supplying the liquid coolant to the evaporator and, at the same time, preparing the chilled coolant by recovering the latent heat of the liquefied propylene and/or propane," the evaporator 3 which is illustrated in Fig. 2 and Fig. 3 may be cited. When the pipe line for liquefied propylene in the evaporator 3 is supplied with propylene and, at the same time, the liquid coolant is introduced into the interior of the evaporator 3 incorporating the pipe line, heat exchange proceeds between the propylene and the liquid coolant, with the result that the liquid coolant will be cooled by recovering the latent heat generated in consequence of the gasification of the liquefied propylene and consequently enabled to prepare the chilled coolant. The means for preparing the chilled coolant may be provided with the liquid coolant supplying system 1 for storing the liquid coolant, the heater for heating the liquid coolant to the appropriate

may be additionally provided with the liquid coolant transfer pump 28.

This invention effectively utilizes the latent heat which used to be discarded and consequently enables the energy consumption for cooling to be reduced. It, therefore, obviates the necessity for the power line distribution heretofore indispensable in the operation of various heat exchangers.

10 Examples

Now, this invention will be described more specifically below with reference to working examples.

(Example of production of catalyst)

15 In 150 liters of purified water kept heated and stirred, 100 kg of ammonium molybdate and 6.3 kg of ammonium paratungstate were dissolved. To this solution, an aqueous nitrate solution separately prepared by mixing a solution of 55 kg of cobalt nitrate in 100 liters of purified water, 20 a solution of 19 kg of ferric nitrate in 30 liters of purified water, and a solution of 22.9 kg of bismuth nitrate in 30 liters of purified water incorporating 6 liters of concentrated nitric acid was added dropwise. A solution of 14.2 kg of an aqueous 20 mass % silica sol solution and 0.29 25 kg of potassium nitrate in 15 liters of purified water was continuously added thereto. The suspension thus obtained was heated under stirring, evaporated to dryness, and then dried and pulverized. The produced powder was formed in cylinders 5 mm in diameter and calcined under stream of air at 460°C 30 for six hours to afford a catalyst.

(Example 1)

Acrylic acid was produced by using the apparatus for the production illustrated in Fig. 2. First, from the liquid coolant supplying system 1, water which temperature had been adjusted at 30°C by the liquid coolant thermocontroller 23 was supplied in the flow amount of 70 m³/h to the propylene evaporator 3. Liquefied propylene was supplied meantime at the flow amount of 2500 kg/h to the evaporator 3. The vaporizing pressure of the propylene was 0.6 MPa in gauge pressure. As a result, the chilled coolant having the temperature of 27°C was obtained at the flow amount of 70 m³/h. This chilled coolant was further cooled up to 20°C by using refrigerator 13.

This chilled coolant was circulated to and used in absorbing solvent cooler 8, the condenser 10 attached to the solvent separating column, and the condenser 11 attached to the refining column. The used chilled coolant was introduced to the liquid coolant supplying system 1 and reused as liquid coolant. To the circulation cooler 9 attached to the acrylic acid absorbing column, the chilled coolant having the temperature of 27°C was directly supplied through the line 2.

The reaction tubes in the reactor 4 for catalytic gas phase oxidation were packed with 5.6 m³ of the catalyst (composition: Mo: 12, Bi: 1, Fe: 1, Co: 4, W: 0.5, Si: 1, and K: 0.06 in molar ratio) obtained in the example of production. The reactant gas was composed of 6 vol. % of propylene, 12 vol. % of oxygen, 9 vol. % of steam, and 73 vol. % of nitrogen, etc. The reactant gas was introduced so as to acquire the retention time of 2 seconds on the catalyst. The results were as shown below.

Fluctuation of pressure of propylene gas in the evaporator: Not more than ±2 kPa

Power consumption in refrigerator: 467 MWh/Y

Efficiency of absorption of acrylic acid immediately after start of operation: 99.5%

Efficiency of absorption of acrylic acid after six months' operation: 99.4%

Efficiency of absorption of acrylic acid after one year's operation: 99.2%

Number of suspensions of apparatus due to the polymerization: 0/year

(Comparative Example 1)

Acrylic acid was produced by following the procedure of Example 1 while supplying saturated steam having the temperature of 120°C and the gauge pressure of 0.1 MPa at the flow volume of 390 kg/h instead of the liquid coolant to the propylene evaporator 3, discarding the steam after heat exchange without deriving any chilled coolant therefrom, and obtaining the heat source for heat exchange in the attached devices wholly from the liquid coolant supplying system and the refrigerator. The vapor pressure of propylene was 1.0 MPa in gauge pressure. The results were as shown below.

Fluctuation of pressure of propylene gas in evaporator: Not more than $\pm 10 - 20$ kPa

Power consumption in refrigerator: 937 MWh/Y

Efficiency of absorption of acrylic acid immediately after start of operation: 99.5%

Efficiency of absorption of acrylic acid after six months' operation: 97.2%

Efficiency of absorption of acrylic acid after one year's operation: Unmeasurable

Number of suspensions of apparatus due to the polymerization: 2/year

(Example 2)

Acrylic acid was produced by using the apparatus for the production illustrated in Fig. 3. First, from the liquid coolant supplying system 1, brine (aqueous 30 mass % ethylene glycol solution) had been adjusted to the temperature of 12°C by the liquid coolant thermocontroller 23 was supplied at the flow amount of 120 m³/h to the propylene evaporator 3. The vapor pressure of propylene was 0.4 MPa in gauge pressure. Liquefied propylene was supplied meantime at the flow amount of 4200 kg/h to the evaporator 3. As a result, the chilled coolant having the temperature of 9°C was obtained at the flow amount of 120 m³/h. This chilled coolant was combined with the liquid coolant from the liquid coolant supplying system 1 and then cooled up to 8°C by using refrigerator 13.

This chilled coolant was circulated to and used in the same manner as in Example 1 while omitting the circulation thereof to the condenser 11 attached to the refining column.

The reaction tubes in the reactor 4 for catalytic gas phase oxidation were packed with 5.6 m³ of the catalyst (composition: Mo: 12, Bi: 1, Fe: 1, Co: 4, W: 0.5, Si: 1, and K: 0.06 in molar ratio) obtained in the example of production. The reactant gas was composed of 6 vol. % of propylene, 12 vol. % of oxygen, 9 vol. % of steam, and 73 vol. % of nitrogen, etc. The reactant gas was introduced so as to acquire the retention time of 2 seconds on the catalyst. The results were as shown below.

Fluctuation of pressure of propylene gas in evaporator:
Not more than ±2 kPa

Power consumption in refrigerator: 3063 MWh/Y

Efficiency of absorption of acrylic acid immediately after start of operation: 99.6%

Efficiency of absorption of acrylic acid after six months' operation: 99.4%

Efficiency of absorption of acrylic acid after one year's operation: 99.3%

5 Number of suspensions of apparatus due to the polymerization: 0/year

(Comparative Example 2)

10 Acrylic acid was produced by following the procedure of Example 2 while supplying saturated steam having the temperature of 120°C and the gauge pressure of 0.1 MPa at the flow volume of 650 kg/h instead of the liquid coolant to the propylene evaporator 3, discarding the steam after heat exchange without deriving any chilled coolant therefrom, and
15 obtaining the heat source for heat exchange in the attached devices wholly from the liquid coolant supplying system and the refrigerator. The vapor pressure of propylene was 1.0 MPa in gauge pressure. The results were as shown below.

20 Fluctuation of pressure of propylene gas in evaporator: Not more than $\pm 10 - 20$ kPa

Power consumption in refrigerator: 3835 MWh/Y

Efficiency of absorption of acrylic acid immediately after start of operation: 99.4%

25 Efficiency of absorption of acrylic acid after six months' operation: 97.0%

Efficiency of absorption of acrylic acid after one year's operation: Unmeasurable

30 Number of suspensions of apparatus due to the polymerization: 2/year

(Results)

(1) In Comparative Example 1, the fluctuation of the

Claims:

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5 (1) A method for the production of acrylic acid or acrolein, characterized by gasifying liquefied propylene and/or propane introduced into an evaporator by supplying a liquid coolant to the evaporator and, at the same time, preparing a chilled coolant by recovering the latent heat of the liquefied propylene and/or propane, subjecting the resultant gasified propane and/or propylene to a catalytic gas phase oxidation reaction thereby preparing a gas containing acrylic acid or
10 acrolein, and using said chilled coolant in heat exchangers attached to the apparatus for the production acrylic acid or acrolein.

15 (2) A method according to claim 1, wherein the temperature of said liquid coolant supplied to said evaporator is in the range of 0 - 50°C.

(3) A method according to claim 1, wherein said liquid coolant is water or brine.

20 (4) A method according to claims 1, where in said heat exchanger is at least one member selected from the group consisting of an absorbing solvent cooler and a circulation cooler attached to the acrylic acid absorbing column, a condenser attached to the solvent separating column, and a condenser attached to the acrylic acid refining column.

25 (5) A method according to claims 1, wherein said chilled coolant is used in heat exchangers attached to a methacrylic acid and/or (meth)acrylic esters plant.

(6) A method according to claims 1, wherein the chilled coolant which has been used in said heat exchangers is reused as a liquid coolant.

30 (7) A method according to claims 1, wherein the amount of said liquefied propylene and/or propane to be gasified is adjusted by controlling the amount of said liquid coolant

into the evaporator.

(8) An apparatus for the production of acrylic acid or acrolein, comprising means for gasifying liquefied propylene and/or propane introduced into an evaporator by supplying
5 a liquid coolant to said evaporator and, at the same time, preparing a chilled coolant by recovering the latent heat of the liquefied propylene and/or propane, means for subjecting the resultant gasified propylene and/or propane to a catalytic gas phase oxidation reaction thereby preparing
10 a gas containing acrylic acid or acrolein, and means for using said chilled coolant in heat exchangers attached to the apparatus for the production of acrylic acid or acrolein.

(9) An apparatus according to claim 8, wherein said means for preparing said chilled coolant includes means for
15 adjusting the temperature of said liquid coolant or means for adjusting the flow amount thereof.

(10) An apparatus according to claim 8, which further comprises means for circulating the chilled coolant used in
20 said heat exchangers to said means for preparing said chilled coolant.

Abstract

Chilled coolant is prepared by liquid coolant utilizing the latent heat generated by a gasification of liquefied propylene, for example, and this chilled coolant is used in heat exchangers which are used in a process for production of acrylic acid or acrolein. This method allows effective utilization of the latent heat which used to be discarded and permits a reduction of energy consumption of cooling required separately in the step for production. By recovering the chilled coolant with the liquid coolant, it makes possible to stabilize the gasification of propylene, etc. and consequently stabilize the production of acrylic acid. This invention consists of providing the method for the production of acrylic acid, etc. and the apparatus which make effective use of the latent heat generated in the steps of production.

FIG.1

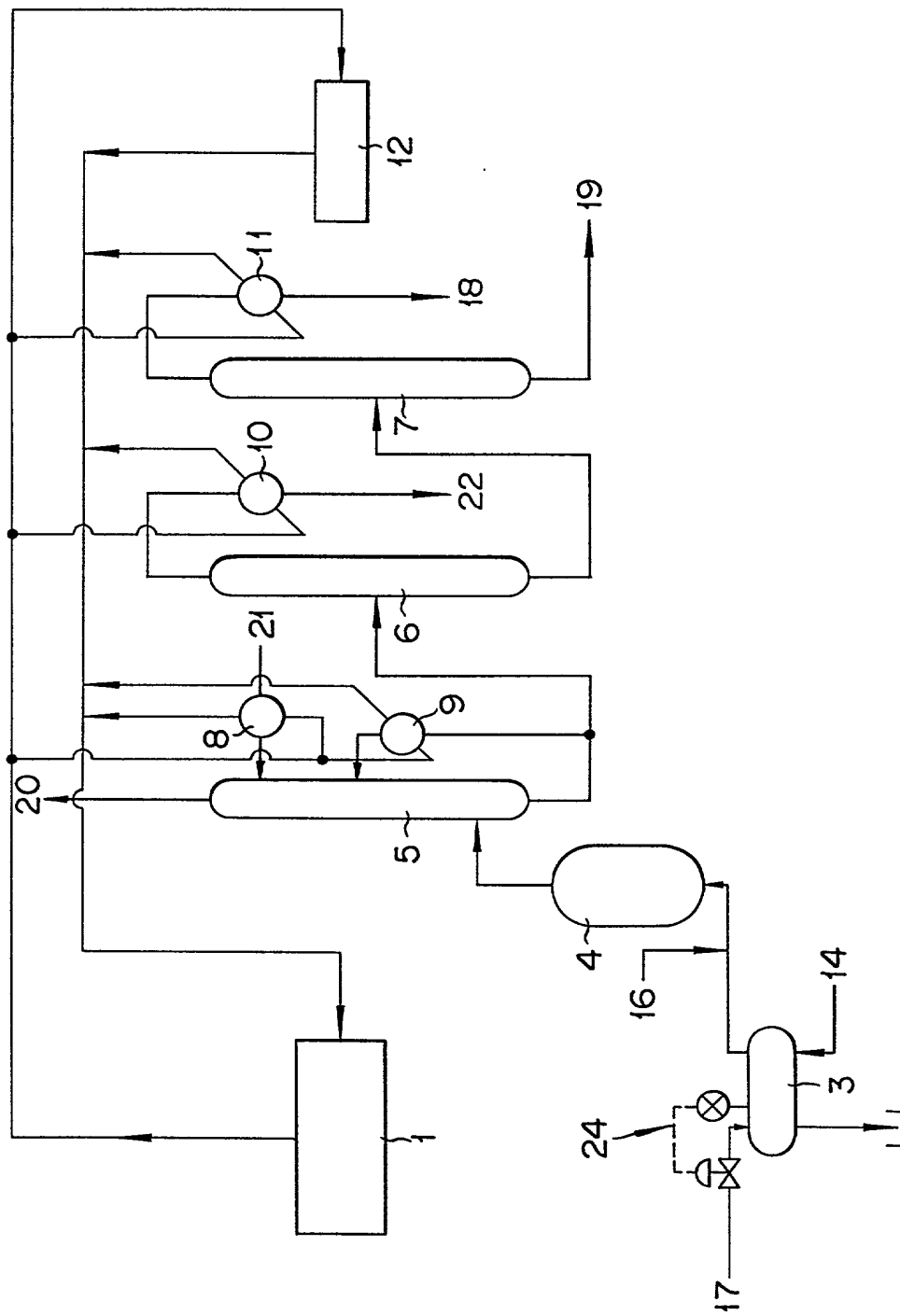


FIG. 2

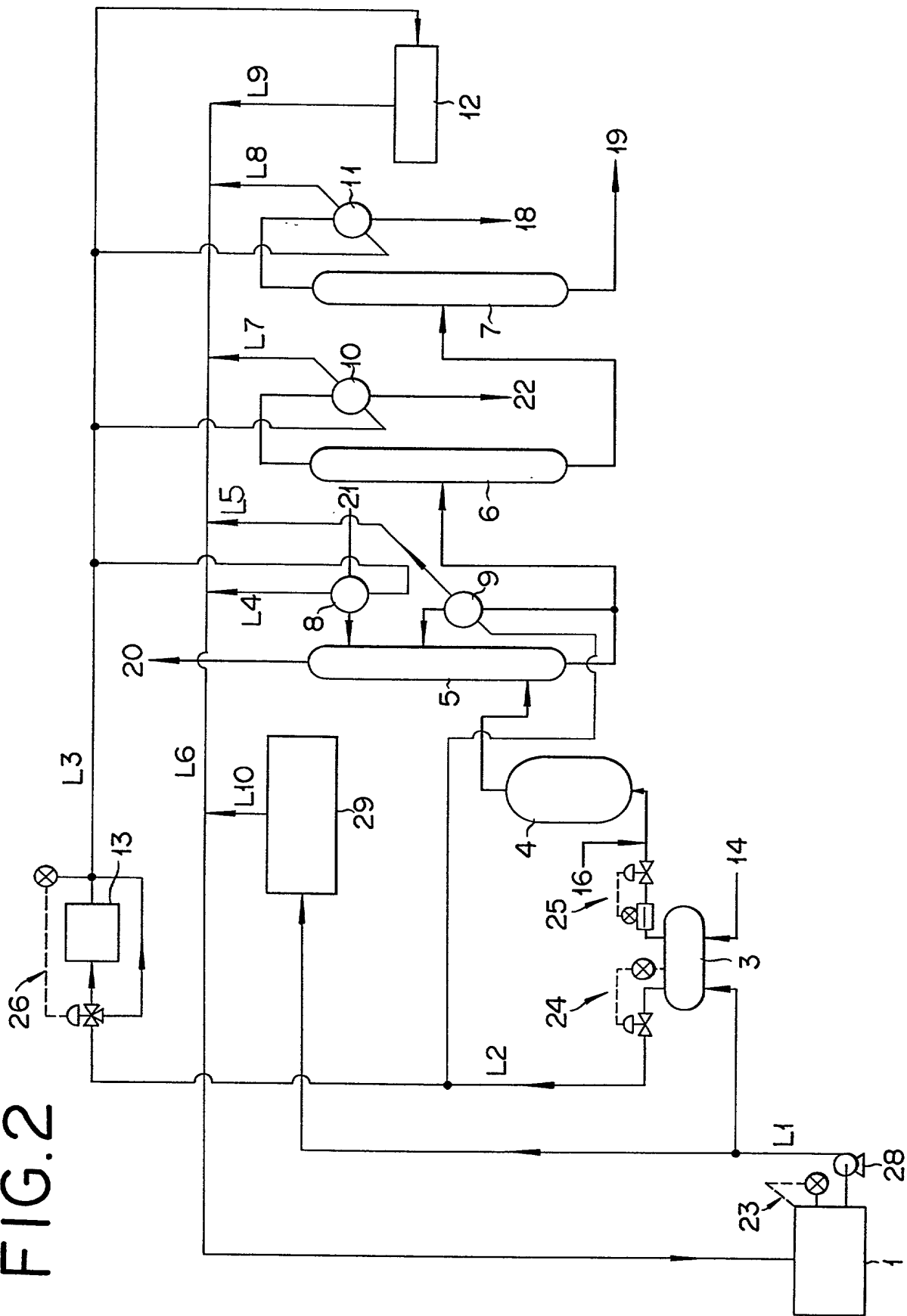
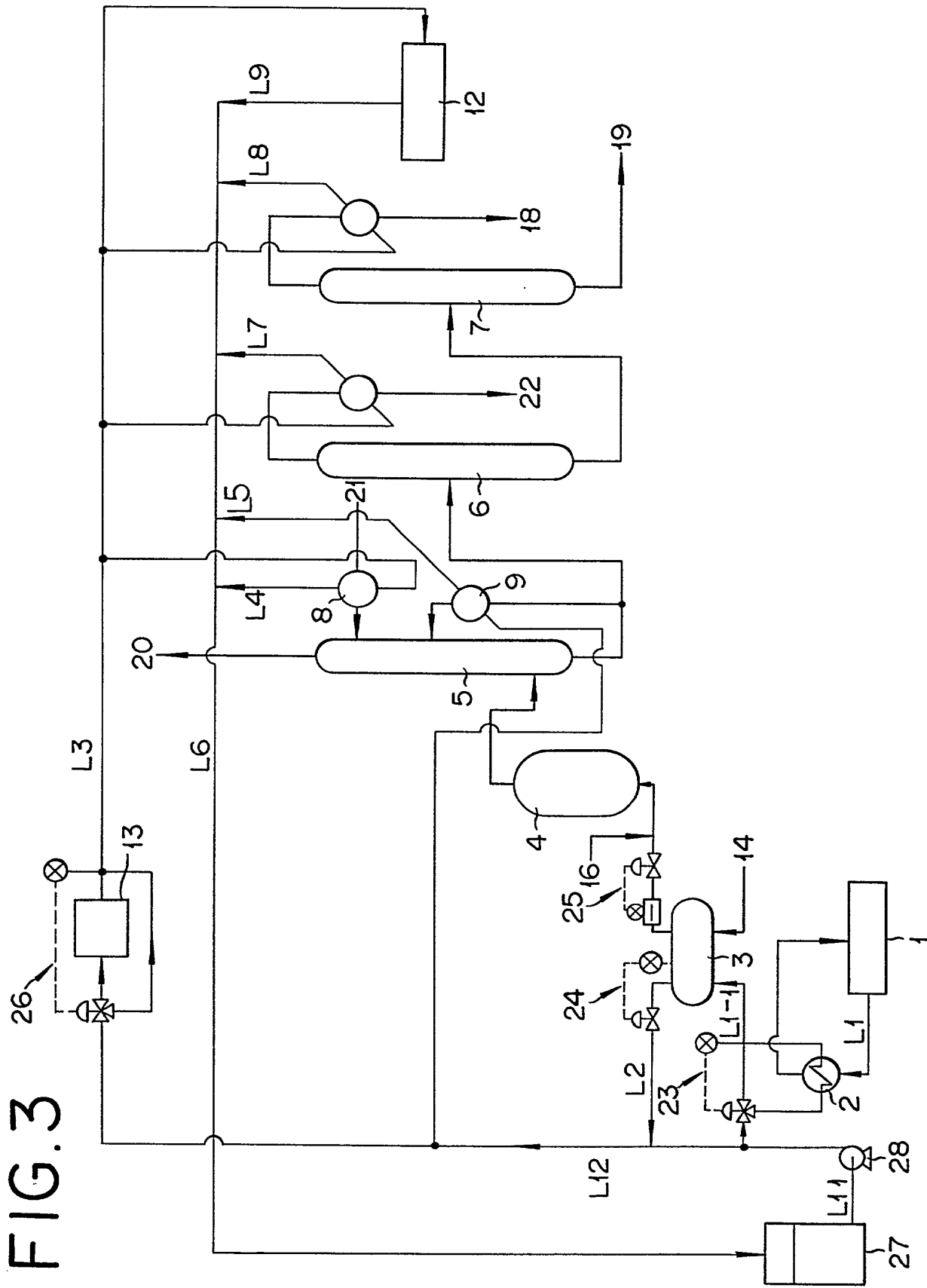


FIG. 3



English Language Declaration

My residence, post office address, and citizenship are as stated below next to my name,

the specification of which

☒ is attached hereto

I hereby state that I have reviewed and understand the contents of the above-identified specification, including the claims, as amended by any amendment referred to above.

I acknowledge the duty to disclose to the United States Patent and Trademark Office all information known to me to be material to patentability as defined in Title 37, Code of Federal Regulations, Section 1.56.

I hereby claim foreign priority benefits under Title 35, United States Code, Section 119(a)-(d) or Section 365(b) of any foreign application(s) for patent or inventor's certificate, or Section 365(a) of any PCT international application which designated at least one country other than the United States, listed below and have also identified below, by checking the box, any foreign application for patent or inventor's certificate or PCT international application having a filing date before that of the application on which priority is claimed.

Prior Foreign Application Number(s)	Country	Foreign Filing Date (MM/DD/YYYY)	Priority Not Claimed	Certified Copy Attached??	
11-315914	Japan	November 5, 1999		YES	NO
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			<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
			<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>
			<input type="checkbox"/>	<input type="checkbox"/>	<input type="checkbox"/>

☐ Additional foreign application numbers are listed on a supplemental priority data sheet PTO/SB/02B attached hereto.

I hereby claim the benefit under 35 U.S.C. 119(e) of any United States provisional application(s) listed below.

Application Number(s)	Filing Date (MM/DD/YYYY)	
		<input type="checkbox"/> Additional provisional application numbers are listed on a supplemental priority data sheet PTO/SB/02B attached hereto.

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